Copyright © by Holt, Rinehart and Winston. All rights reserved.

## SECTION 2-1 REVIEW

## **COMPOSITION OF MATTER**

## **VOCABULARY REVIEW** Define the following terms.

1.	atom								
2.	neutro	neutron							
3.	compound								
4.	covalent bond								
5.	ion								
MU	LTIPL	E CHOICE Write the	correct letter in th	e blank.					
	<b>1.</b> The atomic number of carbon is 6. Therefore, the number of protons in a carbon atom eq								
		<b>a.</b> 3.	<b>b.</b> 6.	<b>c.</b> 7.	<b>d.</b> 12.				
	<b>2.</b> One of the kinds of particles found in the nucleus of an atom is the								
		<b>a.</b> proton.	<b>b.</b> electron.	<b>c.</b> ion.	<b>d.</b> boron.				
	3.	s in an atom's second							
		<b>a.</b> 2.	<b>b.</b> 4.	<b>c.</b> 6.	<b>d.</b> 8.				
<b>4.</b> Of the following elements, the one that is most likely to form ionic bonds is									
		a. hydrogen.	<b>b.</b> carbon.	<b>c.</b> sodium.	<b>d.</b> oxygen.				
		a. water.	<b>b.</b> hydrogen gas.	<b>c.</b> oxygen gas.	<b>d.</b> chloride ion.				

9

Nall	ne			Class	Date			
SHO		<b>R</b> Answer the	questions in the	e space provided.				
1.	What is the difference between mass and weight?							
2.	Identify the elements and the number of atoms of each element in each of the following compounds							
	BO <sub>2</sub>			KCl				
	C <sub>6</sub> H <sub>12</sub> O <sub>6</sub>			NH <sub>3</sub>				
3.	How many pairs of electrons do the two oxygen atoms in an oxygen molecule share with each							
	other? Explain your answer.							
STF	RUCTURES A	ND FUNCTION	<b>VS</b> Use the figur	e to answer the fo	llowing questions.			
	The diagram l of hydrogen is energy level c	he diagram below shows bonding of a hydrogen atom with a chlorine atom. The atomic numbe f hydrogen is 1. The atomic number of chlorine is 17. The orbitals corresponding to the third nergy level can hold up to 8 electrons.						
			0 —	→ 6000000000000000000000000000000000000				
		Cl	Н	HCI				
1.	What kind of	What kind of bond is formed between hydrogen and chlorine atoms?						